



Vanadium-Bromine Redox Flow Battery

**Flow Batterie Kolloquium in Karlsruhe am 27. September
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Stationary Energy Storage – Why Now?

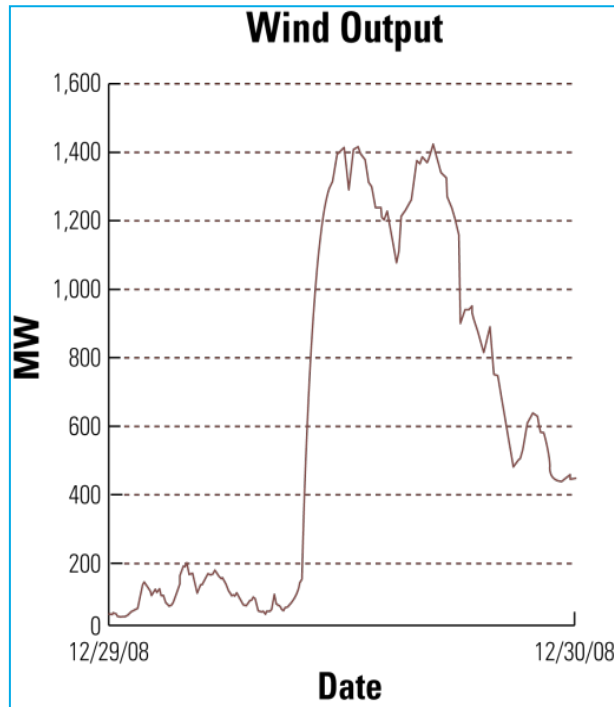


In the past few years the demand for large-scale energy storage has increased for several applications

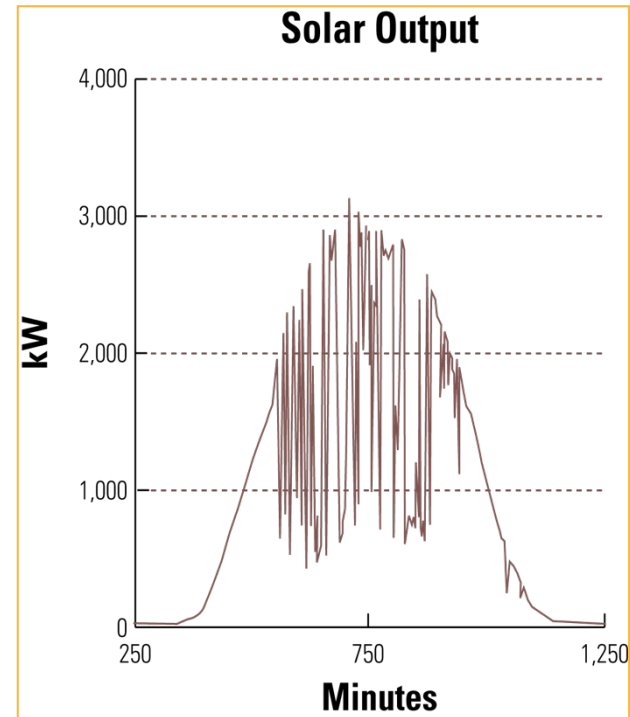
- **Renewables integration**
- Ancillary services
- Arbitrage
- Grid asset optimization
- T&D deferral
- Telecommunications – substitute for diesel

Renewables Integration

Wind & Solar Variability Problem

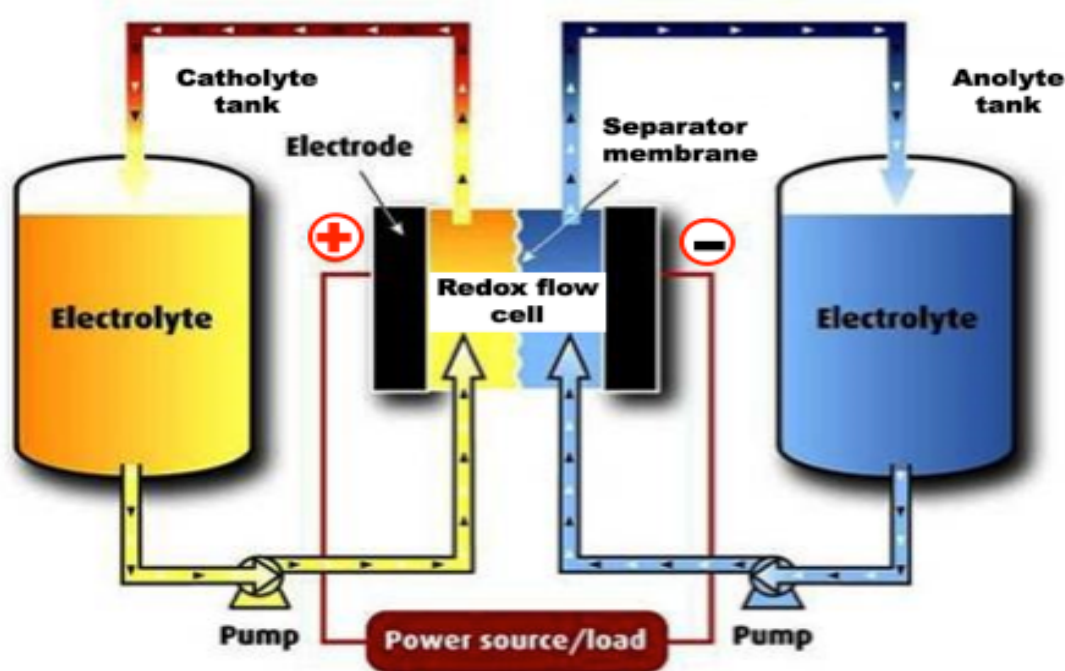


Oregon Wind Farm
10 - 100% in 1 hour



Arizona Solar Farm
10 - 80% in 5 min

Why **Redox** Flow Batteries?



- Separation of POWER and ENERGY
- More POWER = larger stack
- More ENERGY = larger tank
- Competitive energy efficiency

- Highly durable: 10,000+ cycles
- Safe, non-flammable liquid
- Lowest CapEx & OpEx
- Large application space

Product Development Approach



- Start with a proven chemistry to lower risk: **vanadium redox**
- Engineer a next-gen OEM platform: **quantum improvement**
- Identify key barriers preventing full commercialization
- Utilize open innovation approach to secure best IP
- Leverage the best expertise wherever it is
- Develop a multi-generational pipeline of product improvements

Building a Superior Product



| CURRENT PROBLEM SET | WATTJOULE IMPROVEMENT |
|--|--|
| Electrochemical stacks are large and use expensive materials | New breakthroughs now allow major reductions in stack size |
| Electrolyte energy density is low and requires large tanks | We have technology that increases the energy we can store in every liter |
| System needs refrigerated cooling thereby increasing costs and lowering efficiency | We can now eliminate AC and chiller equipment with our platform |
| Lifetime system efficiency needs improvement | We now add major efficiency improvements with no added cost |
| Need for costly, high purity active materials like vanadium | We have a pathway to utilizing less and then no vanadium over time |
| Relatively high cost vs. attractive economics | All of the above improvements translate to significantly lower cost |

Quantum Improvement Factors



Pathway to Better VRB Metrics

| Key Metric | SOA ¹ | Gen 1 | Gen 2 V-Br | Core Benefit |
|---------------------------------|------------------|-------|------------|---------------------------|
| Converter Stack Power Density | 1X | 6X | 7X | Lower material cost |
| Electrolyte Energy Density | 1X | 2X | 3X | Less liquid required |
| Electrolyte Temperature Range | 1X | 3X | 3.2X | No active cooling needed |
| Roundtrip Efficiency | 1X | 1.1X | 1.2X | Lower life cycle cost |
| Vanadium Cost Reduction | 1X | 1.4X | 2.2X | Lower vanadium cost |
| DC System Capital Cost (\$/kWh) | 600 | 200 | 150 | Significantly lower CapEx |

ElectriStor™ ES10 Test System



2kW, 10kWh Engineering Prototype II, DC Only



1,000-fold increase in power and energy from 2014-2016

So...What's Next?



- We believe that the improvements made in the chemistry, materials and design of our Gen 1 all-vanadium redox flow battery have pushed this system nearly to its maximum performance and minimum cost limits.
- Further improvements in our RFB platform will require a change in the basic system chemistry.
- With financial and technical support from ICL, we have chosen the vanadium-bromine redox flow battery for further development.

V-Br Redox Flow Battery

Performance

- Electrolyte energy density of ≥ 50 Wh/kg
- Operating electrode current density of ≥ 200 mA/cm²
- Maximum power density of ≥ 1000 mW/cm²
- Standard operating temperature of 45°C
- Round-trip DC electrical efficiency of 80%

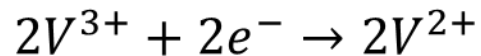
Cost

- \$150/kWh for DC energy storage system

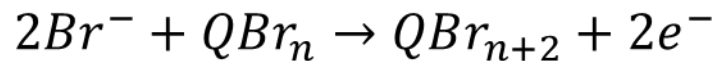
V-Br Redox Flow Battery

Electrode reactions for charge:

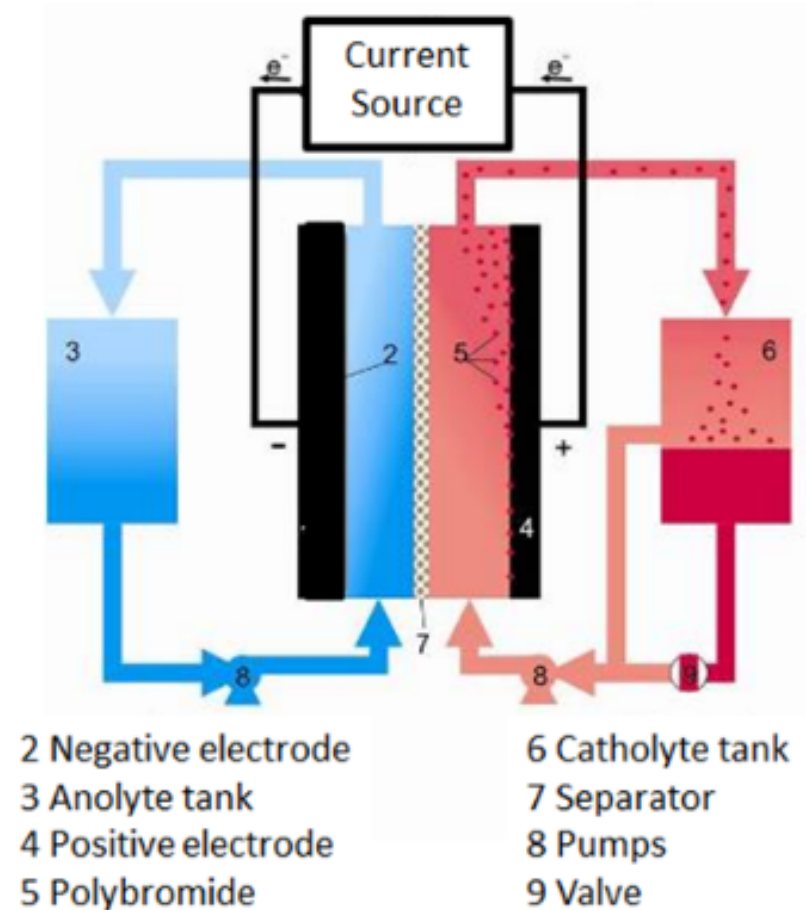
Negative electrode reaction:



Positive electrode reaction:



Insoluble bromine oil falls to the bottom of the catholyte tank

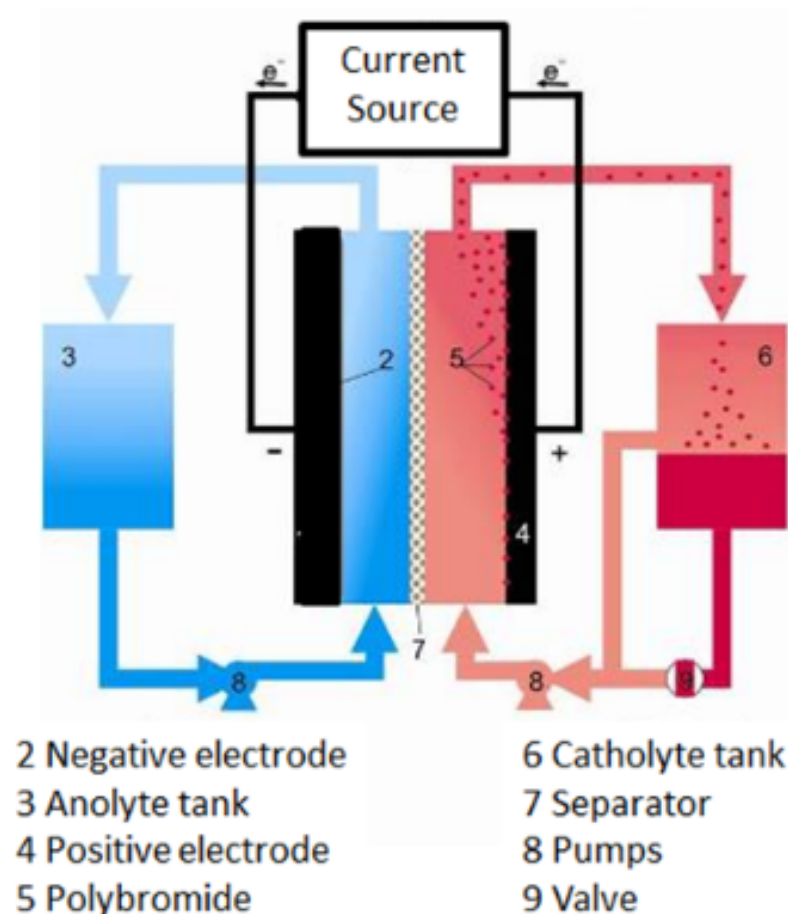


V-Br Redox Flow Battery

Advantages Over All-Vanadium



- Decrease amount of Vanadium by nearly 50%
- Increased electrolyte energy density
- Technology demonstrated in lab-scale hardware
- Utilizes Gen 1 high-power density cell technology
- Proprietary complexing agent provides multiple system benefits
- Strong WattJoule IP position



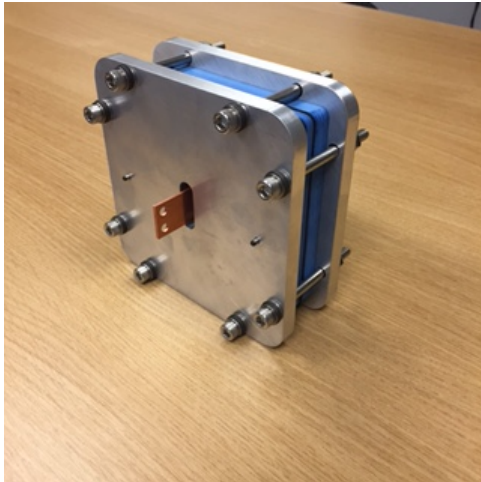
V-Br Redox Flow Battery

Advantages Over Other Chemistries

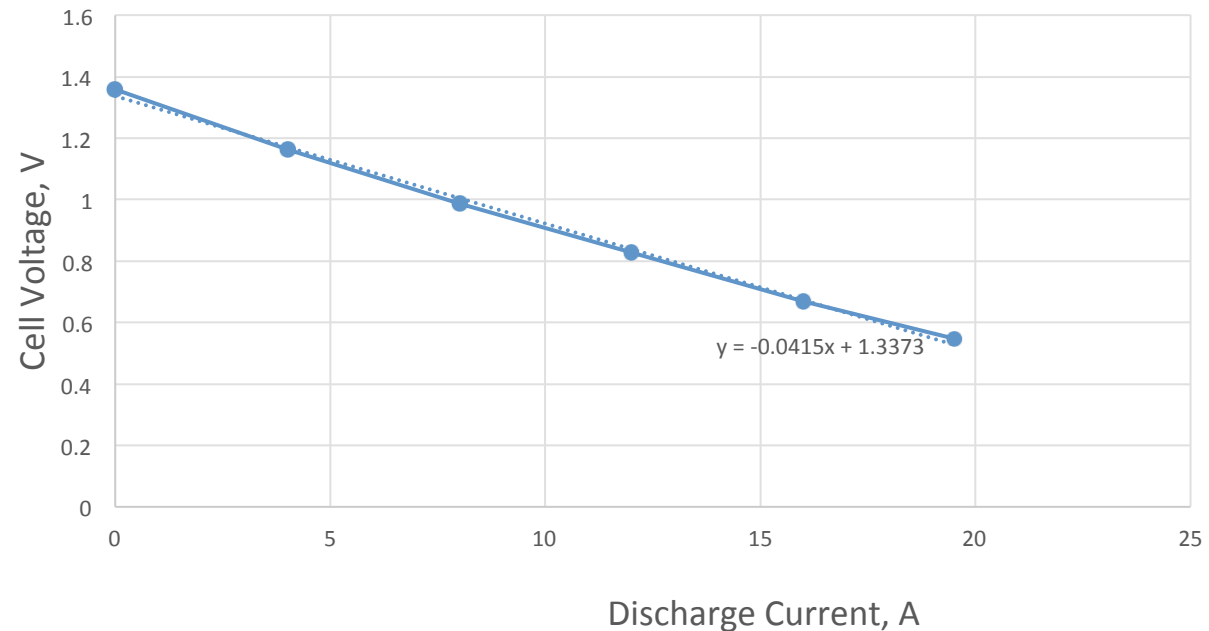


| ALTERNATIVE FLOW CHEMISTRIES | VANADIUM BROMINE |
|--|--|
| Hydrogen bromine requires large high pressure tanks to store flammable and explosive gaseous hydrogen, and needs expensive catalyst that degrades over time | Requires no catalyst and all the energy is safely stored in liquid form. Electrolyte contains over 60% water and cannot burn or explode. |
| Zinc bromine has dendrite problems on electrodes that require stripping and have durability issues. Power and energy coupled since hybrid flow. Low power density. | True redox flow battery that requires no plating and therefore has no dendrite problems. Power and energy capability completely uncoupled. |
| Iron chromium has a significant hydrogen and chlorine gassing problem under normal operation and has low energy density. | Virtually no gassing potential due to electrochemical operating mode. Much higher energy density can be achieved. |
| Vanadium-vanadium requires large stacks and tanks and the higher cost of vanadium, also has a limited temperature range. | Much higher power and energy density can be achieved while cutting vanadium use by 50%. Temperature range not an issue. |

V-Br Test Results



Initial Results of Polarization Test on Gen 2
At 45°C in 25-cm2 Cell



V-Br Test Results



Gen 2 Power Capability in 25-cm² cell at 45°C Vanadium Concentration 2.0 Mol/L

| Membrane | Pmax, mW/cm ² | Specific Resistance, Ωcm ² |
|--------------------------|--------------------------|---------------------------------------|
| Fluorinated Ion-Exchange | 431 | 1.0375 |

Power Performance Similar to Gen 1 in First Experiments

V-Br Test Results



Typical UNSW Cycling with Bromine Complexation Current Density only 10 mA/cm²

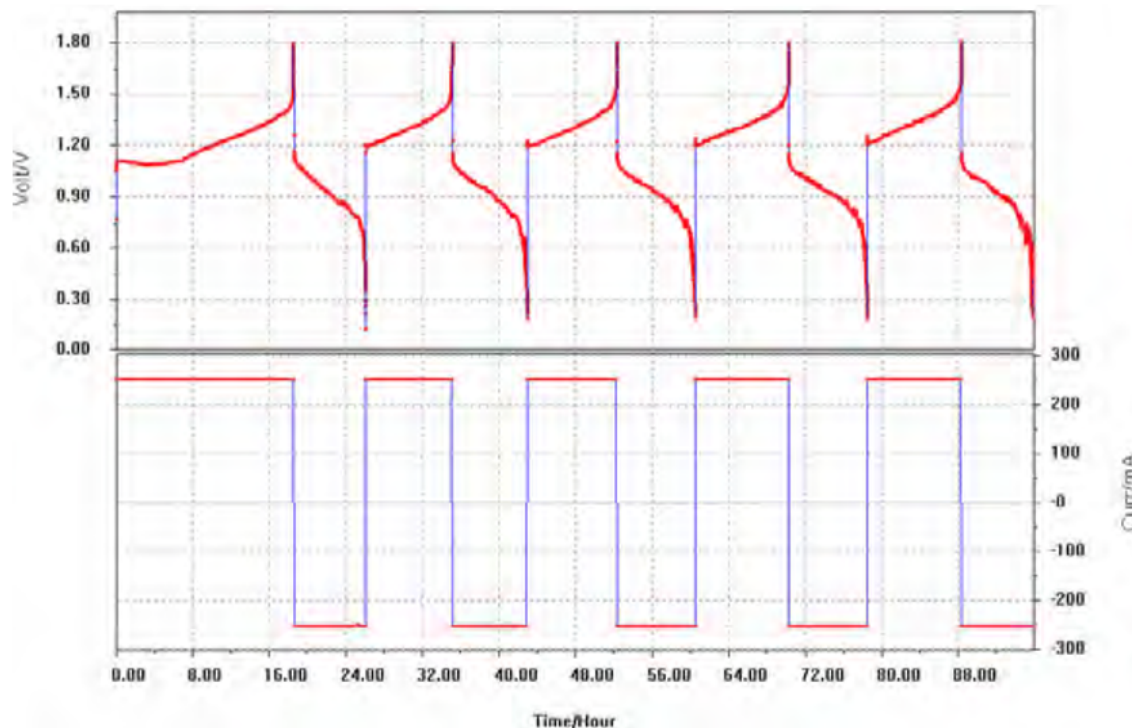
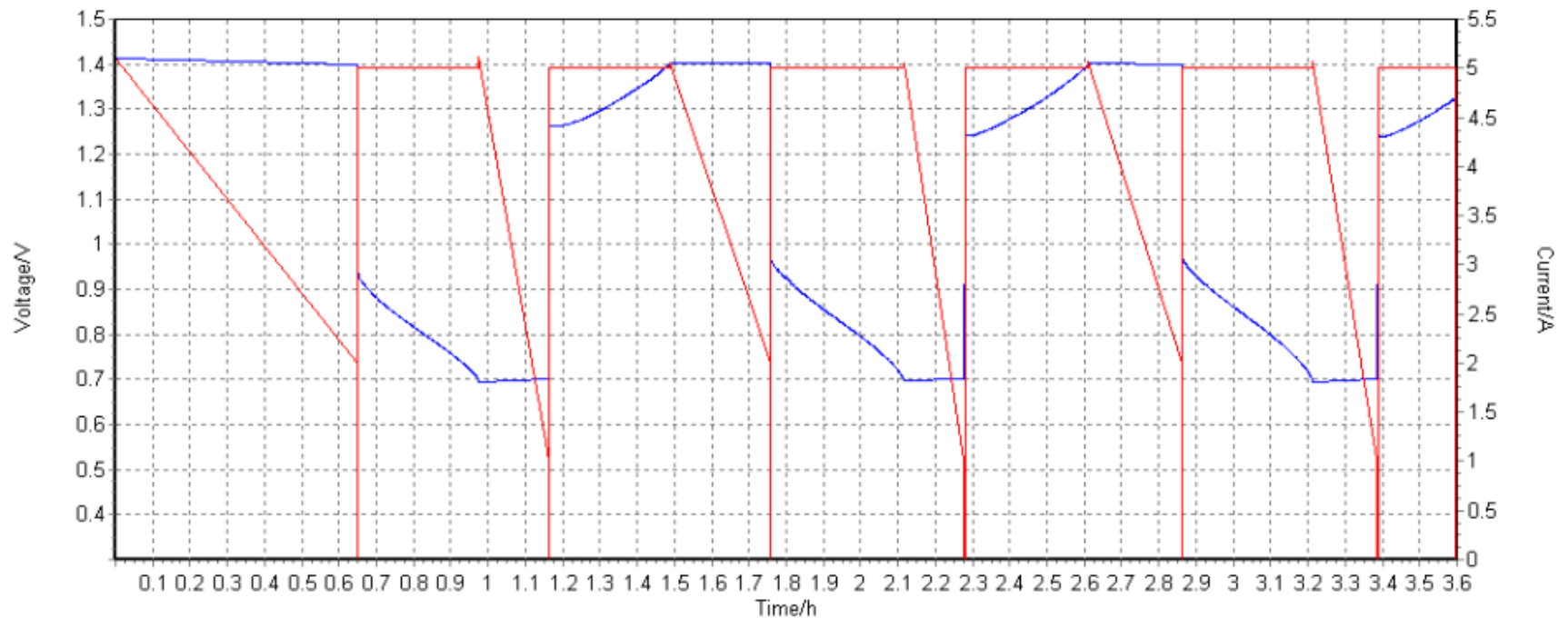


Figure H.23 Charge / discharge cycles of 2 M $V^{3.7+}$, 0.19 M MEM, 0.56 M MEP, 6.1 M HBr, 1.2 M HCl using ChiNafion membrane at 25°C (CY060929.cel; 50ml electrolytes; (+ve) 27°C (-ve) 26°C)

V-Br Test Results



WattJoule Cycling Results at 200 mA/cm² 2M V at 45°C



— V of 0523 Gen2 cycling.004
— I of 0523 Gen2 cycling.004

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